

Quiz 7
Chemical Engineering Thermodynamics
February 23, 2017

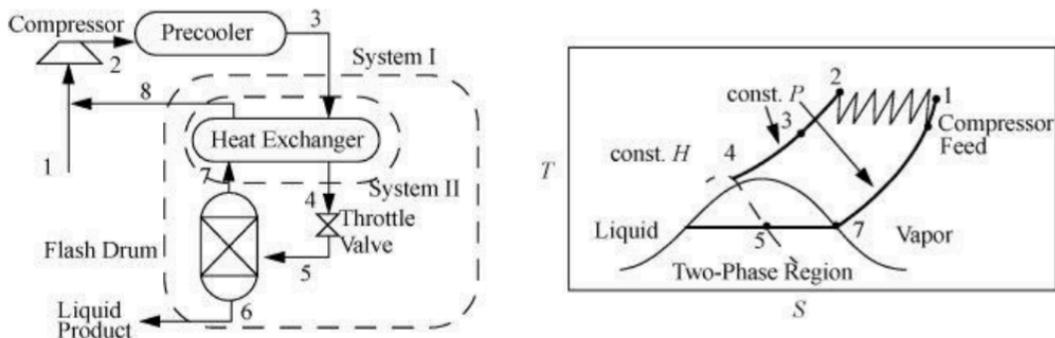


Figure 5.12. Linde liquefaction process schematic. The system boundaries shown on the left

- Figure 5.12 shows the Linde liquefaction process. Sketch a P/H diagram showing the eight streams of the Linde process. (show 2 and 2')
- Make a table for each stream indicating the state (sat. vap. etc.), T , P , H , S , η_0 , q , dm/dt (mass flow rate). Assuming that you have the P/H diagram and that you are given the pressure and temperature for streams 3 and 8. Which streams can be immediately solved for H and how. Fill in as much as you can from the initial information including indicating which streams have the same flow rate, P , T , H , S and how you would use η_0 .
- Can streams 1 and 2 be initially ignored? Why?
- The two dashed lines in Figure 5.12 indicate two local balances that need to be considered to solve this design. Show how the System I balance in mass and in energy can yield q for stream 5.
- Show how q can yield H_5 .
- Show how H for stream 4 can be resolved from an energy balance on system II.

c) Streams 1 & 2 can be ignored since they only provide T & P for stream 3.

$$d) \quad \dot{m}_3 = \dot{m}_6 + \dot{m}_8 \quad \dot{m}_6 = \dot{m}_3 - \dot{m}_8 \quad (1)$$

$$H_3 \dot{m}_3 = H_6 \dot{m}_6 + H_8 \dot{m}_8 \quad (2)$$

$$g = \frac{\dot{m}_8}{\dot{m}_3}$$

using (1) in (2)

$$H_3 \dot{m}_3 = H_6 (\dot{m}_3 - \dot{m}_8) + H_8 \dot{m}_8$$

$$H_3 = H_6 \left(1 - \frac{\dot{m}_8}{\dot{m}_3}\right) + H_8 \left(\frac{\dot{m}_8}{\dot{m}_3}\right)$$

$$(H_3 - H_6) = \frac{\dot{m}_8}{\dot{m}_3} (H_8 - H_6)$$

$$g = \frac{\dot{m}_8}{\dot{m}_3} = \frac{(H_3 - H_6)}{(H_8 - H_6)}$$

$$e) \quad H_5 = H_{L, P_0} + g (H_{U, P_0} - H_{L, P_0})$$

$$f) \quad H_4 \dot{m}_3 - H_3 \dot{m}_3 = H_8 \dot{m}_8 - H_7 \dot{m}_8$$

$$\frac{\dot{m}_8}{\dot{m}_3} = g \quad \therefore H_4 = g (H_8 - H_7) + H_3$$